

Diffusion of Gases Demonstration Worksheet

Data Table

Part A.	Procedure –
	Observations –
Part B.	Procedure –
	Observations –
Part C.	Procedure –
	Observations –
Part D.	Procedure –
	Observations –

Discussion Questions

1. What is gas diffusion?
2. Describe the evidence from Part D that hydrochloric acid molecules diffuse at a slower rate than ammonia molecules.
3. Ammonia molecules move at a different rate than hydrochloric acid molecules because ammonia molecules have a lower mass (17 g/mole) than hydrochloric acid (36.5 g/mole). Explain in terms of the kinetic energy of gas molecules at the same temperature.