FLINN SCIENTIFIC

Name_____

Hydrolysis of Salts Worksheet

Results Table

Petri Dish	Salt	Solution Color	pH	Acid, Base, or Neutral
1	Aluminum chloride			
2	Ammonium chloride			
3	Sodium chloride			
4	Sodium bicarbonate			
5	Sodium phosphate			

Discussion Questions

1. Explain what happened to the salts in the water and what caused the acid–base properties of the solutions.

2. Salt hydrolysis can be described in two chemical equations, the first showing the dissociation of the salt, and the second net equation showing the production of H^+ or OH^- ions. Write the two equations for each salt in this demonstration. If neither H^+ nor OH^- ions are produced, write "no reaction" for the second equation.

   a. $AlCl_3 \cdot 6H_2O(s) \rightarrow$

   b. $NH_4Cl(s) \rightarrow$

   c. $NaCl(s) \rightarrow$

   d. $NaHCO_3(s) \rightarrow$

   e. $Na_3PO_4(s) \rightarrow$