

Data Table

Boiling point range — ethyl acetate (°C)	
Boiling point — beginning of distillation (°C)	
Boiling point — end of distillation (°C)	
Mass of Erlenmeyer flask (g)	
Mass of Erlenmeyer flask plus ethyl acetate (g)	
Mass of ethyl acetate (g)	
Theoretical yield of ethyl acetate (g)	
Percent yield of ethyl acetate (%)	
Fragrance	

Post-Laboratory Review Questions

1. In Pre-Lab question 6c, the theoretical yield of ethyl acetate, in grams, was calculated. Enter this value in the Data Table.

2. Use the mass of the ethyl acetate collected from the distillation and the theoretical yield to calculate the percent yield of ethyl acetate. Record this value in the Data Table.

3. Look up the literature value of the boiling point of ethyl acetate in a reference book, such as the *Merck Index* or the *CRC Handbook of Chemistry and Physics*. Compare the experimental and literature values of the boiling point.