

# Foiled Again — Aluminum Loses to Copper Demonstration Worksheet

## Results Table

Part	Chemicals Used	Observations
Part 1		
Part 2		
Part 3		
Part 4		

## Discussion Questions

- Write a balanced chemical equation for the reaction between aluminum foil and cupric chloride.
  
- Aluminum metal is oxidized into aluminum(III), while the copper(II) is reduced. Write three balanced chemical equations:
  - For the oxidation half-reaction
  
  - For the reduction half-reaction
  
  - For the overall balanced equation
  
- Propose an explanation for why there was no reaction with the aluminum by the cupric sulfate alone or by the sodium chloride alone, but there was when the cupric sulfate and the sodium chloride were mixed together.