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Hydrolysis of Salts Worksheet

Results Table

Petri Dish	Salt	Solution Color	рН	Acid, Base, or Neutral
1	Aluminum chloride			
2	Ammonium chloride			
3	Sodium chloride			
4	Sodium bicarbonate			
5	Sodium phosphate			

Discussion Questions

1. Explain what happened to the salts in the water and what caused the acid-base properties of the solutions.

- 2. Salt hydrolysis can be described in two chemical equations, the first showing the dissociation of the salt, and the second net equation showing the production of H⁺ or OH⁻ ions. Write the two equations for each salt in this demonstration. If neither H⁺ nor OH⁻ ions are produced, write "no reaction" for the second equation.
 - $a. \text{ AlCl}_3 \cdot 6\text{H}_2\text{O}(\text{s}) \rightarrow$
 - b. $NH_4Cl(s) \rightarrow$
 - c. NaCl(s) \rightarrow
 - *d*. NaHCO₃(s) \rightarrow
 - e. Na₃PO₄(s) \rightarrow

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