

Data Table

Mass of crucible + M_2CO_3	g
Mass of crucible + M_2CO_3 (dried) (1st weighing)	g
Mass of crucible + M_2CO_3 (dried) (2nd weighing)	g
Mass of crucible	g
Mass of M_2CO_3	g
Mass of filter paper + $CaCO_3$	g
Mass of filter paper + $CaCO_3$ (2nd weighing)	g
Mass of filter paper	g
Mass of $CaCO_3$	g
Moles of $CaCO_3$	mol
Molar mass of M_2CO_3	g/mol
Identity of M_2CO_3	
Percent error	

Calculations and Post Lab Analyses *(Show all work.)*

1. Calculate the moles of precipitated calcium carbonate, $CaCO_3$. Enter this value in the Data Table.

2. Calculate the molar mass of the unknown carbonate. Enter this value in the Data Table.

3. From the calculated molar mass, identify the unknown. Calculate the percent error in the molar mass value. Enter both values in the Data Table.

4. Review the procedure and list possible sources of errors that would cause the molar mass of the unknown to be (a) too high or (b) too low.